

Amendments to the Specification:

Please replace the paragraph at page 1, directly under the title with the following paragraph:

--This application is a Continuation of U.S. Patent Application Serial No. 09/882,156, filed June 15, 2001; which is a Divisional of U.S. Patent Application Serial No. 09/129,208, filed August 4, 1998, now U.S. Patent No. 6,251,348; which is a Continuation-in-Part of U.S. Patent Application Serial No. 08/757,228 filed Nov. 27, 1996, now U.S. Pat. No. 5,792,438; which claims the benefit of U.S. Provisional Application Serial No. 60/024,239 filed Aug. 20, 1996.—

Please replace the paragraph beginning at page 6 line 17 and carrying over to page 7 with the following rewritten paragraph:

--Metal oxides, such as iron an zinc oxide, have an electronegative potential, meaning the potential is on the active or anodic end of the Emf series, with the active end relating to metals which tend to corrode. Noble metals, copper for example, have an electropositive potential, meaning the potential is on the noble or cathodic end of the Emf series. Cathodic metals do not readily corrode. The Emf series is a listing of elements according to their standard electrode potential. When two dissimilar metals, a noble metal and an active metal, are combined a galvanic cell is formed, which will result in galvanic corrosion. Corrosion of a metal is increased because of the current caused in a galvanic cell, so that as the corrosion rate is increased so is the reactivity of the metal. In particular, when copper, for example, is added to iron oxide, for example, a galvanic cell is formed which causes the iron oxide to corrode faster and thus be more reactive with various sulfur species. What this means is that increasing the electro-potential relates to forming a ~~galvanic~~galvanic cell so that corrosion is increased and reactivity with various sulfur species is increased. Most of this information, as well as,

the Emf series were discussed and disclosed in the "Basic Corrosion Course" offered by the National Association of Corrosion Engineers in October of 1978.—

Please replace the paragraph beginning at page 7, line 16 and carrying over to page 8 with the following rewritten paragraph:

--Because the inventive activator effectively raises the rate of reactivity of oxide products, this invention results in the improvement in the use of disposable oxide products for the removal of sulfur compounds from natural gas and vapors, and other hydrocarbon liquids. Thus, the present invention is desirable because it allows for an oxide product that can be contained in a small reactor vessel, results in the removal of sulfur compounds in a short amount of time, the removal of sulfur compounds from unsaturated hydrocarbon fluids, the non-inclusion of air, and the thorough removal of the sulfur compounds from fluids.—

Please replace the paragraph beginning at page 8, line 2 with the following rewritten paragraph:

~~--Fig. 4~~ The figure is an X-ray diffraction reading showing the presence of copper oxide after having hydrogen sulfide pass through an iron oxide product.—

Please replace the paragraph beginning at page 9, line 12 and carrying over to page 10 with the following rewritten paragraph:

--The activator causes increased reactivity in the iron oxide or zinc oxide product, referred to generally throughout as the iron oxide product, because it has a higher electro-potential than the iron oxide product, with the dissimilar electro-potential causing bi-metallic coupling between the activator, copper oxide for example, and the iron oxide product. This bi-metallic coupling results in an increased rate of reaction between the iron oxide product and the sulfur compounds found in fluids, in particular hydrocarbon

fluids. The activator causes the iron oxide to be more reactive by increasing the corrosion rate of the iron oxide, which causes an increased reactivity between the iron oxide product and sulfur compounds. Essentially, the activator causes the iron oxide to react with the sulfur compounds before the activator reacts with the sulfur compounds. More specifically, while copper oxide is known to react quickly with hydrogen sulfide, the reaction between the copper oxide and the hydrogen sulfide essentially takes place after the reaction of the activated iron oxide with the hydrogen sulfide, with the reaction between the copper oxide and the hydrogen sulfide continuing longer than the concentration of the activator accounts for. This is demonstrated in ~~Fig. 4~~ the figure, which shows the presence of copper oxide in an iron oxide bed after having sulfur compounds pass through and react with the iron oxide bed. The presence of copper oxide is shown in ~~Fig. 4~~ the figure by line 6, with ~~Fig. 4~~ the figure being an X-ray diffraction reading taken after the activated iron oxide product had removed hydrogen sulfide from hydrocarbon gas. In particular, ~~Fig. 4~~ the figure shows that the copper oxide activated the iron oxide to react first, as the amount of hydrogen sulfide that was passed through the iron oxide bed was equal to eight (8) times more hydrogen sulfide than would be necessary to exhaust the copper oxide present in the iron oxide bed. Because the copper oxide did not completely react with the hydrogen sulfide this indicates that the iron oxide reacted with the hydrogen sulfide before the copper oxide. In addition to coupling with and activating iron oxides, the activator can be used to activate other oxides. The other oxides, besides iron oxide, are oxides having a lower electro-potential than the activator.—

Please replace the paragraph beginning at page 13, line 19 and carrying over to page 14 with the following rewritten paragraph:

--It should be further noted that the presence of oxygen in the fluid containing sulfur compounds further increases the electro-potential differential between the oxide product and the activator. Thus, even smaller vessels with dramatically shorter contact times are possible for ~~order~~ odor control applications and hydrogen sulfide removal systems

with vapors naturally containing, or with the deliberate addition of, air, which may include oxygen.—